



# Standard Guide for Electrochemical Noise Measurement<sup>1</sup>

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## 1. Scope

1.1 This guide covers the procedure for conducting online corrosion monitoring of metals by the use of the electrochemical noise technique. Within the limitations described, this technique can be used to detect localized corrosion activity and to estimate corrosion rate on a continuous basis without removal of the monitoring probes from the plant or experimental cell.

1.2 This guide presents briefly some generally accepted methods of analyses that are useful in the interpretation of corrosion test results.

1.3 This guide does not cover detailed calculations and methods; rather it covers a range of approaches that have found application in corrosion testing.

1.4 The values stated in SI units are to be regarded as standard. No other units of measurement are included in this standard.

1.5 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety and health practices and determine the applicability of regulatory limitations prior to use.*

## 2. Referenced Documents

### 2.1 ASTM Standards:<sup>2</sup>

- G1 Practice for Preparing, Cleaning, and Evaluating Corrosion Test Specimens
- G3 Practice for Conventions Applicable to Electrochemical Measurements in Corrosion Testing
- G4 Guide for Conducting Corrosion Tests in Field Applications
- G5 Reference Test Method for Making Potentiodynamic Anodic Polarization Measurements

<sup>1</sup> This guide is under the jurisdiction of ASTM Committee G01 on Corrosion of Metals and is the direct responsibility of Subcommittee G01.11 on Electrochemical Measurements in Corrosion Testing.

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<sup>2</sup> For referenced ASTM standards, visit the ASTM website, [www.astm.org](http://www.astm.org), or contact ASTM Customer Service at [service@astm.org](mailto:service@astm.org). For *Annual Book of ASTM Standards* volume information, refer to the standard's Document Summary page on the ASTM website.

- G15 Terminology Relating to Corrosion and Corrosion Testing (Withdrawn 2010)<sup>3</sup>
- G16 Guide for Applying Statistics to Analysis of Corrosion Data
- G31 Guide for Laboratory Immersion Corrosion Testing of Metals
- G46 Guide for Examination and Evaluation of Pitting Corrosion
- G59 Test Method for Conducting Potentiodynamic Polarization Resistance Measurements
- G61 Test Method for Conducting Cyclic Potentiodynamic Polarization Measurements for Localized Corrosion Susceptibility of Iron-, Nickel-, or Cobalt-Based Alloys
- G96 Guide for Online Monitoring of Corrosion in Plant Equipment (Electrical and Electrochemical Methods)
- G102 Practice for Calculation of Corrosion Rates and Related Information from Electrochemical Measurements
- G106 Practice for Verification of Algorithm and Equipment for Electrochemical Impedance Measurements

## 3. Terminology

3.1 *Definitions*—The terminology used herein, if not specifically defined otherwise, shall be in accordance with Terminology G15. Definitions provided herein and not given in Terminology G15 are limited only to this guide.

### 3.2 Definitions of Terms Specific to This Standard:

3.2.1 *coupling current, n*—measured current flowing between two electrodes in an electrolyte coupled by an external circuit.

3.2.2 *current measuring device, n*—device that is capable of measuring the current flow across the electrode/electrolyte interface or the coupling current of a pair of electrodes, usually a zero resistance ammeter (ZRA) or current-to-voltage converter.

3.2.3 *electrochemical current noise measurement, n*—electrochemical noise measurement using an electrochemical current signal.

3.2.4 *electrochemical noise measurement (ENM), n*—technique involving the acquisition and analysis of electrochemical current and potential signals.

<sup>3</sup> The last approved version of this historical standard is referenced on [www.astm.org](http://www.astm.org).

3.2.5 *electrochemical potential noise measurement*, *n*—electrochemical noise measurement using an electrochemical potential signal.

3.2.6 *Fourier transform*, *n*—transformation of a time domain signal into the frequency domain.

3.2.7 *galvanostat*, *n*—device used for automatically maintaining a controlled current between two electrodes.

3.2.8 *noise impedance*,  $|Z_n|$ ,  $[\Omega]$ , *n*—ratio of the amplitude of potential noise to current noise, in the frequency domain, at a specified frequency.

3.2.9 *noise resistance*,  $R_n$ ,  $[\Omega]$ , *n*—standard deviation of potential noise divided by the standard deviation of current noise.

3.2.10 *pit indicator*, *n*—standard deviation of current noise divided by the mean of the coupling current.

3.2.11 *pitting factor*, *n*—standard deviation of the current noise divided by the general corrosion current.

3.2.11.1 *Discussion*—The general corrosion current is normally estimated by a secondary electrochemical means.

3.2.12 *pitting index*, *n*—standard deviation of current noise divided by the root mean square of the coupling current calculated over the same sample period.

3.2.13 *potential measuring device*, *n*—a high impedance digital voltmeter or electrometer used to measure the potential between two electrodes.

3.2.13.1 *Discussion*—Ideally, one of these electrodes is under study and the other is a reference electrode; however, the measurements may be made between two nominally identical electrodes manufactured from the material being studied.

3.2.14 *potentiostat*, *n*—device used for automatically maintaining a controlled voltage difference between an electrode under study and a reference electrode in which a third electrode, the counter (or auxiliary) electrode, is used to supply a current path from the electrode under study back to the potentiostat.

3.2.15 *sample interval*, *n*—time delay between successive electrochemical noise measurements.

3.2.16 *sample period*, *n*—time between the first and last data collection during electrochemical noise measurement.

3.2.17 *time domain analysis*, *n*—direct evaluation of time series data, for example, using statistical descriptions of the data.

3.2.18 *time record*, *n*—dataset obtained over a sample period at a typical sample interval in electrochemical noise measurement.

3.2.19 *zero resistance ammeter (ZRA)*, *n*—electronic device used to measure current without imposing a significant IR drop by maintaining close to 0-V potential difference between the inputs.

## 4. Summary of Guide

4.1 Electrochemical noise measurement is used for monitoring of localized corrosion processes such as pitting (1, 2).<sup>4</sup>

4.2 Electrochemical noise measurement may be used to estimate a general corrosion rate (3).

4.3 Electrochemical noise measurement operates on the principle that fluctuations in potential and current occur as a result of spontaneous changes in the instantaneous corrosion rate (4). The fluctuations may be due to one or more of several phenomena that include: initiation (5) and propagation of localized corrosion (6), Faradaic currents (7), double-layer capacitance discharge, gas bubble formation (8), adsorption/desorption processes, surface coverage (9), diffusion (10), variation of film thickness (11), mobility of charge carrier (12), passivity breakdown (13), and temperature variations (14, 15).

4.4 The noise fluctuations associated with corrosion phenomena can usually be distinguished from thermal (white) noise (caused by thermal effects in which noise power is directly proportional to the measured bandwidth), Johnson noise (produced by the measurement instrumentation), and shot noise (in electrical circuits caused by the quantized nature of the electronic charge) (16-18). However, the electrochemical noise signals generated may have characteristics similar to those stated in the preceding sentence.

4.5 The electrochemical noise method of corrosion measurement may help to evaluate the corrosion mechanism of metals in electrolytes. Its particular advantage is in continuous monitoring without applying any external perturbation.

4.6 *Method A—ZRA-Based Current and Potential Measurement*—Two nominally identical electrodes are coupled through a ZRA, which maintains a 0-V potential difference between them by injecting (measured) current. The potential between the couple and a third (reference) electrode is also measured. The reference electrode may be either a conventional reference electrode such as a saturated calomel electrode (SCE) or simply be a third electrode identical in material to the coupled electrodes (19, 20).

4.7 *Method B—Potentiostatic Current Measurement with Standard Reference Electrode*—Reference Test Method G5 provides practice for making potentiostatic measurements. The working electrode potential is controlled with respect to the reference electrode at a prescribed value. The current measured (flowing between the working (Test 1) and auxiliary or counter (Test 2) electrodes) is that required to maintain potential control (21, 22).

NOTE 1—Noise on the reference electrode will result in a corresponding current noise signal; therefore, the reference electrode needs to be relatively noise free. The potential measurement can only be made across the auxiliary and working electrodes, as the potential between the

<sup>4</sup> The boldface numbers in parentheses refer to the list of references at the end of this standard.

reference and the working is held constant by the potentiostat. The voltage developed across the auxiliary and working electrodes is a function of the current flowing through the cell and the impedance caused by the auxiliary electrode, the working electrode, and the solution resistance.

4.8 *Method C—Galvanostatic Potential Measurement*—An electrode is supplied with current from a galvanostat at a prescribed current value. The potential difference between the electrode and a reference electrode is measured. An auxiliary electrode is used to carry the return current.

4.9 There are several methods by which the electrochemical noise data can be obtained (23-26) and analyzed, and some methods of interpreting the data are given in Appendix X1 (27-35). These analyses are included to aid the individual in understanding the electrochemical noise technique and some of its capabilities. The information is not intended to be all-inclusive.

## 5. Significance and Use

5.1 Use of this guide is intended to provide information on electrochemical noise to monitor corrosion on a continuous basis.

5.2 This guide is intended for conducting electrochemical noise measurements, both in the laboratory and in-service environments (36).

5.3 This technique is useful in systems in which process upsets or other problems can create corrosive conditions. An early warning of corrosive attack can permit remedial action before significant damage occurs to process equipment (37).

5.4 This technique is also useful when inhibitor additions are used to control the corrosion of equipment. The indication of increasing corrosion activity can be used to signal the need for additional inhibitor (38).

5.5 Control of corrosion in process equipment requires knowledge of the rate or mechanism of attack on an ongoing basis. This technique can be used to provide such information in a digital format that is easily transferred to computers for analysis (39).

## 6. Limitations and Interferences

6.1 Results are representative of the probe element (electrode). When first introduced into a system, corrosion rates on a probe element may be different from that of the structure.

6.2 Noise can originate from thermal, electrical, and mechanical factors. Since the interest is only on the noise from Faradaic processes, care should be exercised to minimize noise from other sources.

6.3 Probe elements by their nature are consumable. Hazardous situations may occur if probes are left in service for extended periods beyond their probe life. In some configurations, crevice corrosion can cause damage or leaks at the interface between the element and its sealing surface that can cause false readings.

6.4 Electrical contact between probe elements should be avoided. In certain situations (for example, sour corrosion in the presence of hydrogen sulfide), the corrosion products can

lead to apparent electrical shorting of the probe elements leading to erroneous readings.

## 7. Apparatus

### 7.1 *Electronics:*

7.1.1 The input impedance of the device should be high enough to minimize current drawn from the electrodes, such that the electrodes are not polarized by the measuring device. Practice G106 provides guidelines for verification of algorithm and equipment for electrochemical impedance measurements.

7.1.2 The potential range of the device depends on the maximum potential difference between the two electrodes (typically <1 V).

7.1.3 The potential resolution of the device should be adequate to discriminate the signal to within the required accuracy (typically 10  $\mu$ V or lower).

7.1.4 The device should be capable of maintaining an offset potential between the two electrodes of less than 1 mV.

7.1.5 The frequency response of the device should be flat (within the desired accuracy) across the frequency range of the analysis. The device should have a fast enough response so that signal transients are not distorted. Note that the signal that one is attempting to measure may be below the resolution of the instrument.

NOTE 2—The offset voltage will appear as a current offset in the measurement with no electrodes connected.

7.1.6 The current range depends on the system being measured. The wide dynamic ranges seen in passive-to-active transitions (nA to mA) may require auto-ranging circuits.

7.1.7 The bias current of the device should be within the required accuracy of the measurement. Otherwise it may cause an error in the measured current.

7.1.8 The background noise of the device should be below the electrochemical current or potential noise being measured. High-impedance reference electrode inputs may pick up extraneous noise from the environment and shielding may be required. An independent measurement of the background noise level should be performed.

7.1.9 The requirements in 7.1.1 – 7.1.8 do not include all possible combinations of instrumentation and electrode arrangements. The instrument, cell, and analysis requirements should be determined by the particular test being undertaken.

7.2 *Test Cell*—The test cell should be constructed to allow the following items to be inserted into the solution chamber:

7.2.1 Three identical electrodes, two of which comprise the coupled electrodes and the third electrode acts as a reference. Alternatively, instead of the third identical electrode, a Luggin-Haber capillary with salt bridge connection for a reference electrode may be used.

7.2.2 An inlet and an outlet for air or an inert gas.

7.2.3 A thermometer or thermocouple holder.

7.2.4 The test cell shall be constructed from materials that will not corrode, deteriorate, or otherwise contaminate the solution. Practice G31 provides standard practice for conducting immersion corrosion testing.

7.2.5 One type of suitable cell is described in Reference Test Method G5. Cells are not limited to that design. For example, a 1-L round-bottom flask can be modified for the addition of